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1975 E G SKOLNIK, M KAUFMAN

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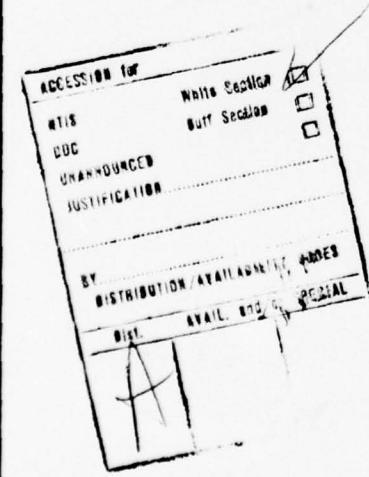
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20. ABSTRACT (Continue on reverse side if necessary and identify by block number) Atomic fluorine was reacted with HCl in a flow system to see if the F + HCl → HF + Cl reaction could be used as a titration for determination of F atom concentration, with an endpoint indicated by the intensity of chlorine atom recombination emission. yields		

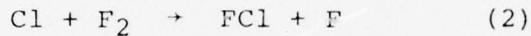
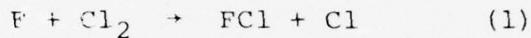
## 20. Abstract (continued)

A sharp intensity maximum of this emission was observed near the endpoint, indicating the rapid removal of chlorine atoms by excess HCl. Since kinetic studies indicated that this removal process was third order overall and first order in [Cl], the Cl + HCl + M → HCl<sub>2</sub> + M reaction was hypothesized. Third order rate constants were determined for M = HCl, Ar, and N<sub>2</sub>. The relative efficiencies were found to be Ar:N<sub>2</sub>:HCl = 1:7.9:19.5, and k<sub>Ar</sub> = 8.2 ± 1.2 × 10<sup>14</sup> cm<sup>6</sup> mole<sup>-2</sup> sec<sup>-1</sup>.

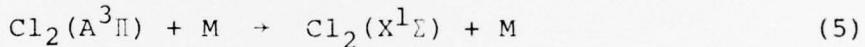
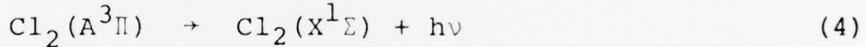
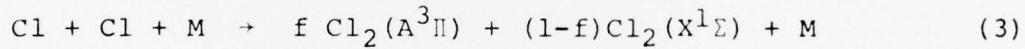


## INTRODUCTION

One method used for the determination of atomic fluorine concentration for kinetic studies has been titration using  $H_2^1$ ,  $NOCl^2$ , and  $Cl_2^{3,4}$  as titrants. In particular, when  $F_2$  is used as the fluorine atom source, molecular chlorine is most useful as a titrant due to the fact that reaction (1) is about four orders of magnitude faster than reaction (2)<sup>3</sup>

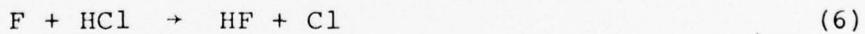


so that virtually no Cl atoms are lost or  $F_2$  molecules consumed by reaction (2). Instead, the Cl atoms recombine homogeneously to form molecular chlorine in both ground ( $X^1\Sigma$ ) and excited ( $A^3\Pi$ ) states<sup>5</sup>. Visible emission from the latter can then be used as a measure of Cl (and thus F) concentration.



Reaction (5), the collisional deactivation of the excited  $Cl_2$  species, is indicated as an alternative to the chemiluminescent decay.

It was the original purpose of this study to determine if the reaction of F atoms with HCl,



followed by reactions (3)-(5) would also be a suitable chemiluminescent titration for atomic fluorine. When it was determined that another process was occurring in the F/HCl system, becoming evident at higher HCl concentrations, the present study turned to the investigation of this other process.

## EXPERIMENTAL

Fluorine atoms were produced by passing F<sub>2</sub>/Ar mixtures through a 2450 MHz microwave discharge produced in an alumina discharge tube by an Evanson cavity excited by a Raytheon Microwave Power Generator. The F/F<sub>2</sub>/Ar mixture then entered the reaction cell, a Teflon coated quartz tube 3.3 cm in diameter and 55 cm in length, where HCl and chlorine and, where necessary, nitrogen were added through a moveable multi-perforated Teflon inlet (Figure 1). Fluorine, HCl and chlorine flows were monitored by standard flow meters. Fluorine (Matheson) flow was controlled by a Nupro stainless steel fine metering valve. Flows of HCl (Linde) and chlorine (Matheson) were controlled by a Granville-Phillips Variable Leak. Argon (Linde) flowed through an Oxytrap column (Alltech) and a P<sub>2</sub>O<sub>5</sub> trap, while nitrogen (Linde) was passed through an Oxysorb column (Alltech). Pressure was monitored by both a tilting McCleod Gauge and Hastings DV-4 and DV-6 Vacuum Gauges. The flow system was pumped through liquid nitrogen and charcoal traps by a Welch Model 1402 Vacuum Pump, having a pumping speed of 160 liters/min. Flow rates varied from 100-500  $\mu\text{mol/sec}$ , depending on the object of the particular experiment.

The emission was detected by a RCA IP28 phototube powered by a Fluke Model 415B High Voltage Power Supply and monitored by a Keithley Model 610C Electrometer. The emission was filtered by a Wrattan 16A filter having a 5160-6200 Å bandpass. Both the photocell and Teflon inlet tube could be moved along the length of the reaction cell, allowing measurements of intensity vs time to be made, while circumventing local wall conditions.

## RESULTS AND DISCUSSION

The reaction of F atoms with Cl<sub>2</sub> in an argon atmosphere was studied at varying Cl<sub>2</sub> concentrations. Cl<sub>2</sub>(A+X) emission intensity was monitored at a position 3 cm downstream from the inlet, so that all initial mixing would be complete. A plot of intensity vs. Cl<sub>2</sub> concentration (Figure 2, solid line) showed a good reproduction of the earlier Ganguli and Kaufman titration<sup>3</sup>, with an easily identified endpoint and emission intensity increasing as the square of the Cl atom concentration up to the endpoint.

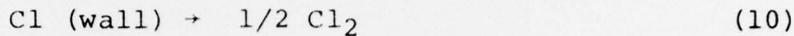
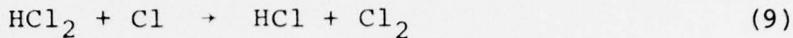
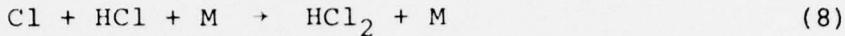
The results of the analogous F + HCl reaction is shown in Figure 2 by the dotted line. While at low concentrations, the square relationship is maintained, there is a sharp and dramatic decrease in intensity at higher HCl concentrations. Although a systematic study was not undertaken, the experiments performed indicated that a comparison of the two titrations, run consecutively using the same F atom concentration produced the same endpoint. Thus HCl would seem to be as effective as Cl<sub>2</sub> as a chemiluminescent titrant for F atoms. In fact, the sharp endpoint may make HCl preferable to Cl<sub>2</sub>.

On the other hand, the results left open the question of what process was responsible for the sharp intensity dropoff. Reaction (6) is well documented as the initial step of the F/HCl reaction<sup>6,7</sup>. If the subsequent reactions are assumed to be reactions (3)-(5), a steady state treatment of Cl<sub>2</sub>(A<sup>3\Pi</sup>) arrives at the result that the intensity is proportional to the square of the Cl atom concentration:

$$I \propto [Cl_2(A^3\Pi)] \propto \frac{fk_3[M][Cl]^2}{(k_4+k_5[M])} \quad (7)$$

If the decay of Cl atoms were second order as in reaction (3), a plot of  $1/I^{1/2}$  vs time would be linear. Figure 3a shows this not the case. Figure 3b shows the linearity of a plot of  $\ln I^{1/2}$  vs time, indicating that the decay of Cl atoms is first order.

The following reaction mechanism for the decay of Cl atoms is thus hypothesized:



The  $\text{HCl}_2$  complex has been discussed<sup>8,9</sup> although there is some controversy as to its existence<sup>10</sup>. This type of mechanism would result in a first order decay of Cl atoms, provided reaction (8) is the rate determining step, and is rapidly followed by reaction (9). A predominant collisional dissociation of  $\text{HCl}_2$  (reverse of reaction (8)) is ruled out, as it would result in a second order Cl atom decay, as can be seen if a steady state approximation is applied to  $\text{HCl}_2$ . The heterogeneous recombination of Cl atoms (reaction (10)) is also included as a well-known<sup>7</sup> first order decay process.

To determine the rate constant  $k_8$  for the rate determining step, pseudo-first order conditions were used in the experiments, with the concentrations of HCl and third body being in large excess (at least twenty times greater) of the F atom (and therefore Cl atom) concentrations. Under these conditions, pseudo-first order rate constants,  $k'$ , were determined as the slope of the  $\ln I^{1/2}$  vs time plots (Figure 3b) and could be related to the third order rate constants as follows:

$$k' = k_8(\text{Ar}) [\text{Ar}] [\text{HCl}] + k_8(\text{HCl}) [\text{HCl}]^2 + k_{10} \quad (11)$$

Only Ar and HCl were present in sufficient concentration to be important third bodies for reaction (8).

The rate constant for argon as a third body,  $k_8(\text{Ar})$  was determined first by finding  $k'$  for a number of different argon concentrations while holding the HCl concentration constant. Plotting  $k'$  vs

[Ar][HCl] gives a slope corresponding to  $k_8(\text{Ar})$ . A typical series of runs for this process is shown plotted in Figure 4, and a summary of all such measurements appears in Table 1. The value found for  $k_8(\text{Ar})$  was  $(8.2 \pm 1.2) \times 10^{14} \text{ cm}^6 \text{ mol}^{-2} \text{ s}^{-1}$ . Once this value was determined, runs were made in which the HCl concentration was varied while argon concentration was held constant. Since  $k_8(\text{Ar})$  is now known, the term  $k_8(\text{Ar}) [\text{Ar}][\text{HCl}]$  can be subtracted from  $k'$  in equation (11). Plots of  $k' - k_8(\text{Ar}) [\text{Ar}][\text{HCl}]$  vs  $[\text{HCl}]^2$  for various HCl concentrations results in a slope corresponding to  $k_8(\text{HCl})$ , with an intercept corresponding to the rate constant for the wall reaction,  $k_{10}$ . A typical series of runs used for the determination of  $k_8(\text{HCl})$  appears in Figure 4 for comparison with  $k_8(\text{Ar})$ , and the summary of all values obtained for  $k_8(\text{HCl})$  and  $k_{10}$  is shown in Table 1. The value for  $k_8(\text{HCl})$  is found to be  $(1.6 \pm 0.4) \times 10^{16} \text{ cm}^6 \text{ mol}^{-2} \text{ s}^{-1}$ , while  $k_{10}$ , although subject to much more scatter, is  $11.7 \pm 3.2 \text{ s}^{-1}$ , corresponding to  $\gamma = (9.1 \pm 2.5) \times 10^{-4}$ . Neither  $k_8(\text{Ar})$  nor  $k_8(\text{HCl})$  show any pressure dependence, indicating that reaction (8) is indeed a third order reaction.

An interesting result is that HCl appears to be about twenty times as efficient a third body as is Ar for the production of  $\text{HCl}_2$ . In order to determine which properties of HCl are responsible for the greatly enhanced efficiency, a number of experiments were performed with  $\text{N}_2$  as an added third body. Although  $\text{N}_2$  is a diatomic molecule, it is inert and nonpolar. A typical series of runs used to determine  $k_8(\text{N}_2)$  is shown in Figure 5. Here both Ar and HCl concentrations are fixed, and  $\text{N}_2$  concentration is varied for different runs. A plot of  $k'$  vs  $[\text{HCl}][\text{N}_2]$  gives a slope equal

to  $k_8(N_2)$ . The value for this rate constant was found to be  $(7.3 \pm 2.0) \times 10^{-15} \text{ cm}^6 \text{ mol}^{-2} \text{ s}^{-1}$ . Thus, the efficiency of  $N_2$  as a third body is nearly nine times that of argon, but only about one-half that of HCl. Therefore while the polarity of HCl appears to be a factor in its efficiency as a third body, the dominant factor appears to be the availability of internal degrees of freedom to remove the energy.

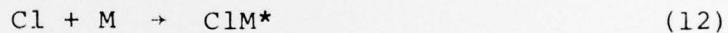
In Table 2, the results of this work is summarized giving both absolute rate constants for  $k_8(M)$ , and rates relative to  $k_8(Ar)$ . In addition, for the sake of comparison, also listed are the relative efficiencies of third bodies in the homogeneous recombination of chlorine atoms (reaction (3)) as found by various authors<sup>12-14</sup>. The nature of the third body employed seems more important in the production of a complex such as  $HCl_2$  than in the production of a simple substance such as  $Cl_2$ .

The mechanism proposed here for the reaction of Cl + HCl lends more support for the existence of  $HCl_2$ . This species has been the subject of some controversy over the past few years. Experimentally, Noble and Pimentel<sup>8</sup> assigned spectral features produced when HCl/ $Cl_2/Ar$  mixtures were passed through a discharge and condensed in a matrix to  $HCl_2$ . Spencer and Glass<sup>9</sup> postulated  $HCl_2$  as a species produced by reaction of Cl atoms with HCl at the chamber walls following the H + HCl reaction. Their measured cross-section for broadening of Cl EPR lines by HCl was also indicative of a strong long range attraction between Cl and HCl.  $HCl_2$  has also been predicted by the BEBO calculations of Truhlar, Olson and Parr<sup>15</sup>, and employed in the trajectory analysis of Thommarson and Berend<sup>16</sup>. In addition, the fact that vibrational relaxation of HCl by Cl atoms has been found to be very rapid<sup>17</sup>, indicates that a  $HCl_2$  intermediate

is likely. The existence of  $\text{HCl}_2$  in the matrix studies, however, has been disputed by Milligan and Jacox<sup>10</sup> and by Wight, Ault and Andrews<sup>18</sup> who claimed the species in question was actually  $\text{HCl}_2^-$ . In addition, Klein and Persky<sup>19</sup> calculated a high energy barrier to the formation of  $\text{HCl}_2$ .

The species present in the current research is certainly not ionic in nature. In addition, the fact that the Cl atom decay is first order, yet reaction (8) is dependent also on HCl concentration and that of the third body indicates the formation in a three-body reaction of an intermediate which could be only  $\text{HCl}_2$ . This is in contrast with the work of Spencer and Glass<sup>9</sup>, where  $\text{HCl}_2$  formation was assumed to largely occur on the walls, in order to explain its varying importance when different wall coatings of their vessel were employed.

Recently, Cohen and Bott reviewed the HCl chemical laser. In their report, they designated the primary recombination mechanism for Cl atoms to be



In comparing the relative efficiencies of the various M's, they stated that qualitatively HCl would be only half as efficient as Ar due to the potential barrier to  $\text{HCl}_2$  formation. The fact that in the current research we find a better than order-of-magnitude increase in HCl third body efficiency over Ar is further proof for the existence of  $\text{HCl}_2$ .

Finally, it should be pointed out that the titration comparison shows the sharp dropoff in intensity only past the endpoint - where HCl is present. This indicates that while Cl atoms will react with HCl to form the  $\text{HCl}_2$  complex, no similar reaction occurs between Cl and HF.

Table 1  
Compilation of Data for  $k_{8(\text{HCl})}$ ,  $k_{8(\text{Ar})}$ , and  $k_{10}$

Series	$k_{8(\text{HCl})}$ ( $\text{cm}^6 \text{ mol}^{-2} \text{ s}^{-1}$ )	Total Pressure (Torr)
1	$1.6 \times 10^{16}$	1.0
9	1.8	3.7
11	1.8	6.05
13	1.1	6.05
21	1.1	5.0
22	2.0	4.1

Mean  $k_{8(\text{HCl})} = (1.6 \pm 0.4) \times 10^{16} \text{ cm}^6 \text{ mol}^{-2} \text{ s}^{-1}$

	$k_{8(\text{Ar})}$ ( $\text{cm}^6 \text{ mol}^{-2} \text{ s}^{-1}$ )	
6	$8.0 \times 10^{14}$	2.95 - 5.8
14	6.6	4.15 - 6.9
15	8.9	4.15 - 6.4
16	9.3	6.05 - 8.0

Mean  $k_{8(\text{Ar})} = (8.2 \pm 1.2) \times 10^{14} \text{ cm}^6 \text{ mol}^{-2} \text{ s}^{-1}$

	$k_{10}$ ( $\text{s}^{-1}$ )	
1	7.7	1.0
5	13.5	3.8
9	8.3	3.7
10	9.9	4.0
11	8.9	6.05
12	12.1	4.3
13	13.0	6.05
21	16.8	5.0
22	14.9	4.1

$$\text{Mean } k_{10} = 11.7 \pm 3.2 \text{ s}^{-1}$$

Table 2  
Summary of Results

M =	Ar	$N_2$	HCl	$Cl_2$	Ref
$Cl + HCl + M:$					
$k_8(M) \times 10^{-15}$ ( $\text{cm}^6 \text{ mol}^{-2} \text{ s}^{-1}$ )	0.82 $(\pm 0.12)$	7.3 $(\pm 2.0)$	16 $(\pm 4)$	--	This Work
Relative Efficiencies	1	8.9	19.5	--	This Work
Comparison with $Cl + Cl + M:$					
Relative Efficiencies	1	--	--	4.6	12
	1	--	--	6.8	13
	1	2	--	--	14

## REFERENCES

1. C. E. Kolb and M. Kaufman, J. Phys. Chem. 76, 942 (1972).
2. K. H. Homann, W. C. Solomon, J. Warnatz, H. Gg. Wagner and C. Zetzsch, Ber. Bunsenges. Phys. Chem. 74, 585 (1970).
3. P. S. Ganguli and M. Kaufman, Chem. Phys. Lett. 25, 221 (1974).
4. P. C. Nordine and D. E. Rosner, J. Chem. Soc. Faraday Trans. 1 72, 1526 (1976).
5. G. Schatz and M. Kaufman, J. Phys. Chem. 76, 3586 (1972).
6. K. L. Komps and J. Wanner, Chem. Phys. Lett. 12, 560 (1972).
7. H. Gg. Wagner, J. Warnatz and C. Zetzsch, Ber. Bunsenges. Phys. Chem. 80, 571 (1976).
8. P. N. Noble and G. C. Pimentel, J. Chem. Phys. 49, 3165 (1968).
9. J. E. Spencer and G. P. Glass, J. Phys. Chem. 79, 2329 (1975).
10. D. E. Milligan and M. E. Jacox, J. Chem. Phys. 53, 2034 (1970).
11. A. C. Lloyd, Internat. J. Chem. Kin. 3, 39 (1971).
12. M. A. A. Clyne and D. H. Stedman, Trans. Faraday Soc. 64, 2698 (1968).
13. L. W. Bader and E. A. Ogryzlo, Nature, 201, 491 (1964).
14. H. Hippler and J. Troe, Internat. J. Chem. Kin. 8, 501 (1976).
15. P. G. Truhlar, P. C. Olsen and C. A. Parr, J. Chem. Phys. 57, 4479 (1972).
16. R. L. Thommarson and G. C. Berend, Internat. J. Chem. Kin. 5, 629 (1973).
17. See for example: N. C. Craig and C. B. Moore, J. Phys. Chem. 75, 1623 (1971); B. A. Ridley and I. W. M. Smith, Chem. Phys. Lett. 9, 457 (1971).
18. C. A. Wight, B. S. Ault and L. Andrews, J. Chem. Phys. 65, 1244 (1976).
19. F. S. Klein and A. Persky, J. Chem. Phys. 59, 2775 (1973).
20. N. Cohen and J. F. Bott in "Handbook of Chemical Lasers," R. W. F. Gross and J. F. Bott, Editors, John Wiley and Sons, New York, N. Y., 1976, Chapter 2.

LIST OF FIGURES

Figure 1            Diagram of the Fast Flow Apparatus

Figure 2            Comparison of Cl<sub>2</sub> and HCl as Titrants for F Atoms  
○ Cl<sub>2</sub>  
□ HCl                Total Pressure: 2.9 Torr

Figure 3a          Decay of Cl Atoms as a Second Order Relationship  
HCl Concentration =  $1.3 \times 10^{-9}$  mol cm<sup>-3</sup>  
Total Pressure = 1 Torr

Figure 3b          Decay of Cl Atoms as a First Order Relationship  
HCl Concentration =  
○  $1.3 \times 10^{-9}$  mol cm<sup>-3</sup>  
□  $2.2 \times 10^{-9}$   
△  $3.7 \times 10^{-9}$   
Total Pressure = 1 Torr

Figure 4            Determination of k<sub>8(M)</sub>  
○ M = Ar, HCl Concentration =  $7.7 \times 10^{-9}$  mol cm<sup>-3</sup>  
□ M = HCl, Ar Concentration +  $208 \times 10^{-9}$  mol cm<sup>-3</sup>

Figure 5            Determination of k<sub>8(N<sub>2</sub>)</sub>  
Ar Concentration =  $275.6 \times 10^{-9}$  mol cm<sup>-3</sup>  
HCl Concentration =  $8.9 \times 10^{-9}$  mol cm<sup>-3</sup>

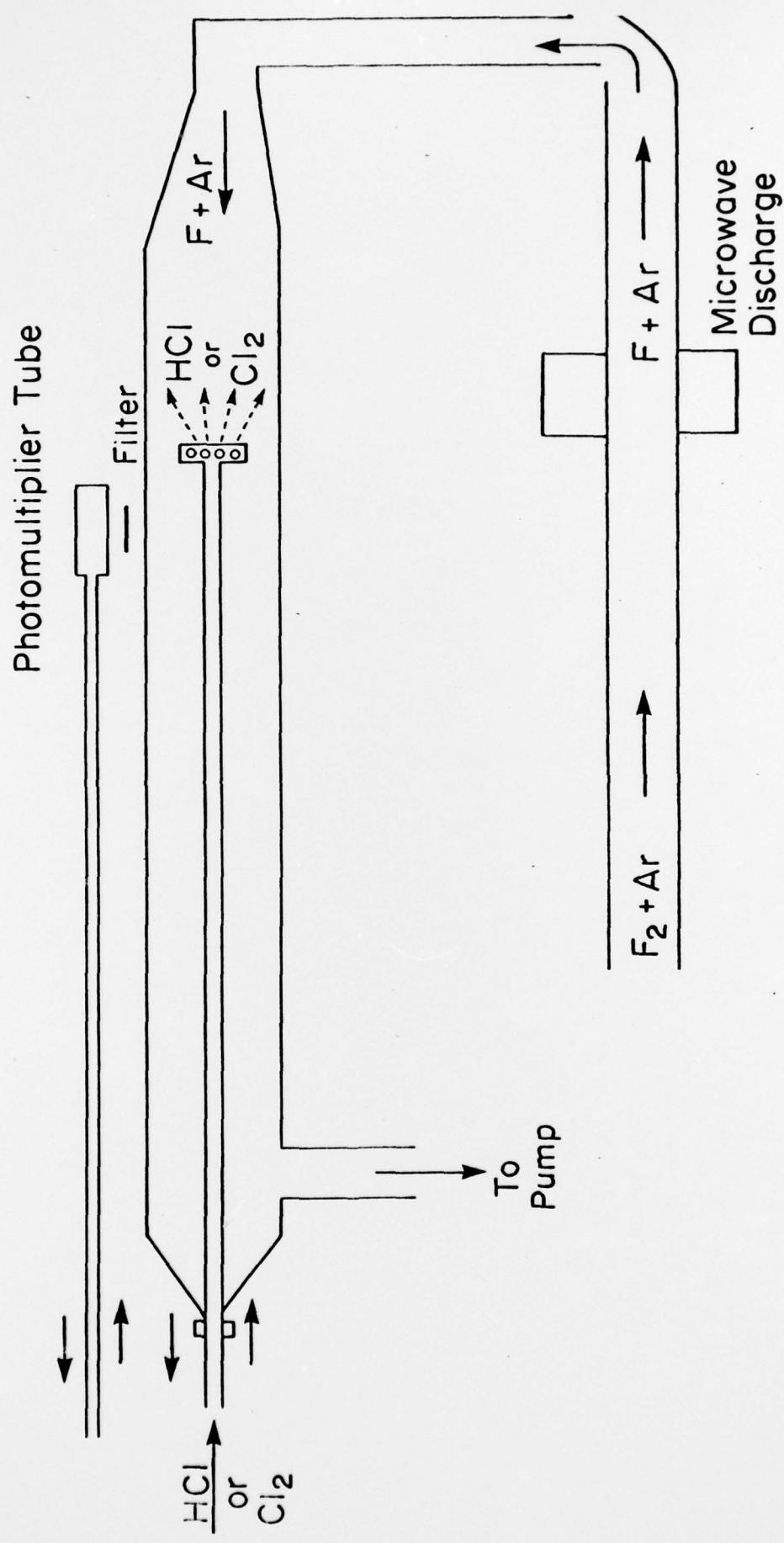


Figure 1

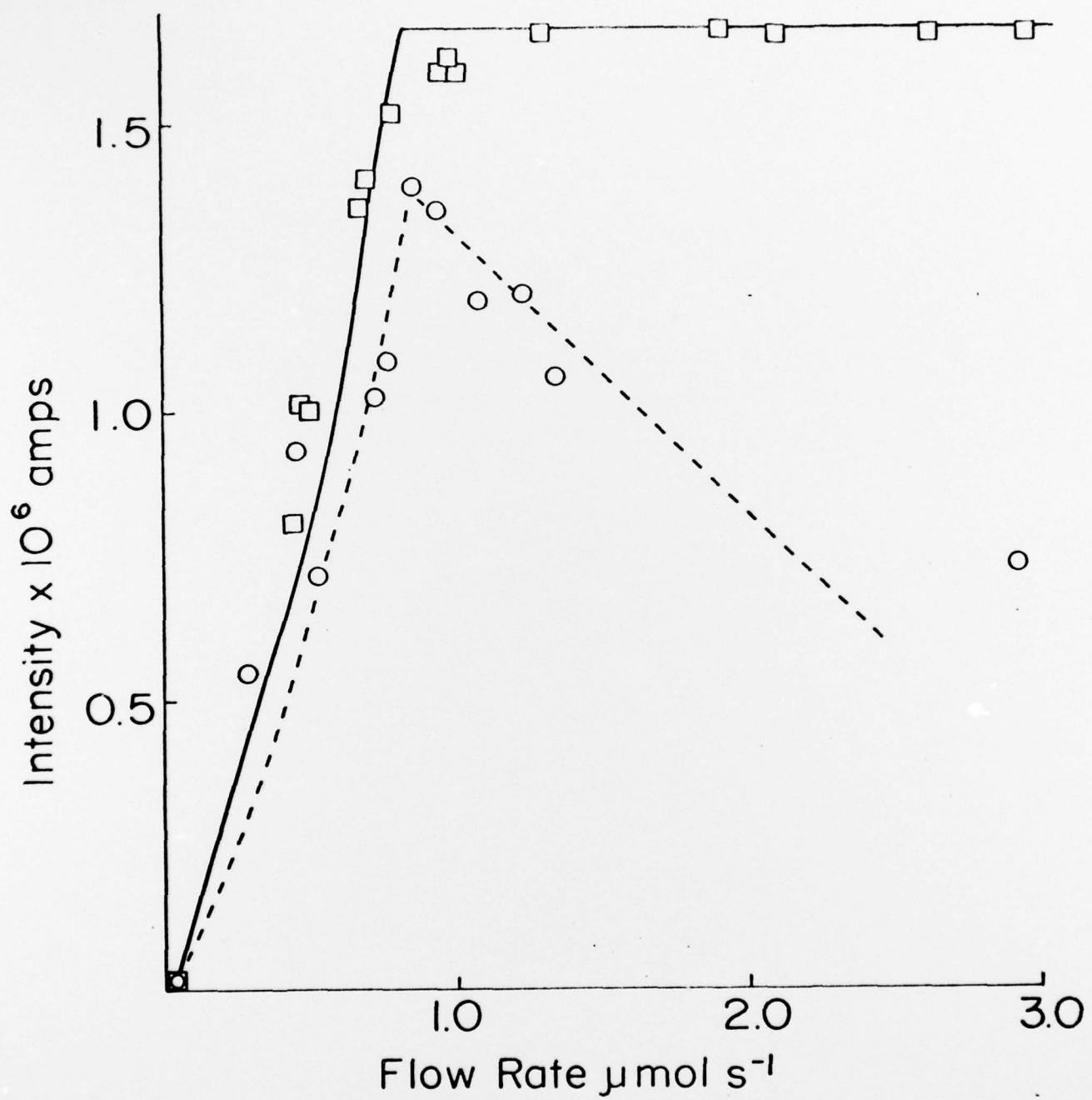


Figure 2

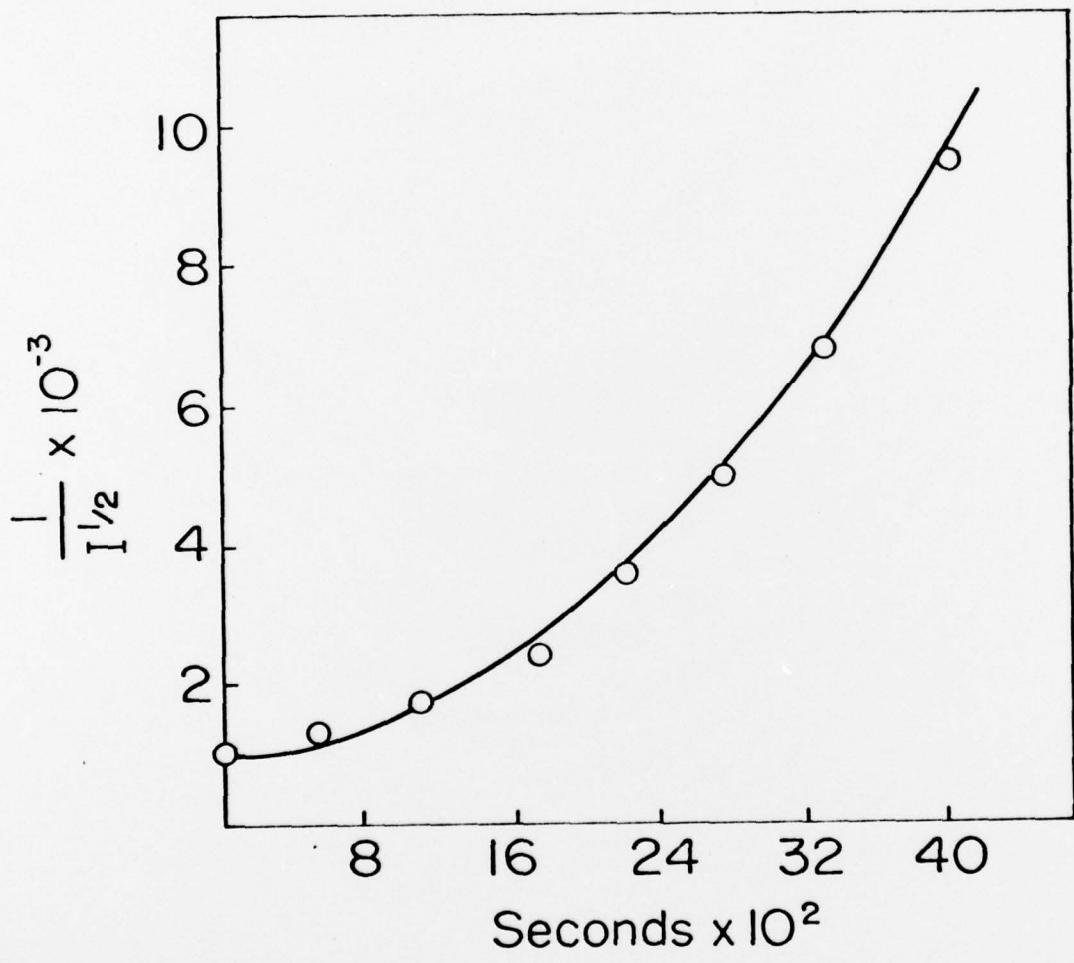


Figure 3a

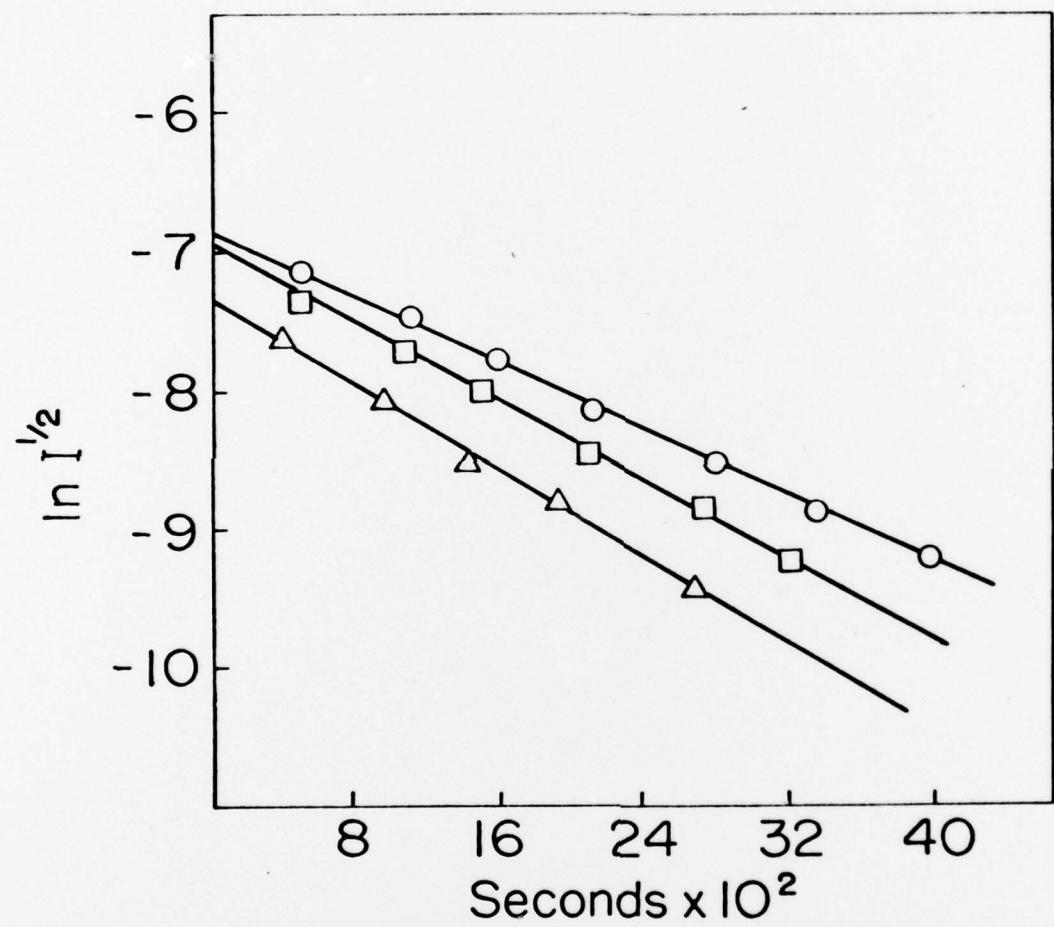


Figure 3b

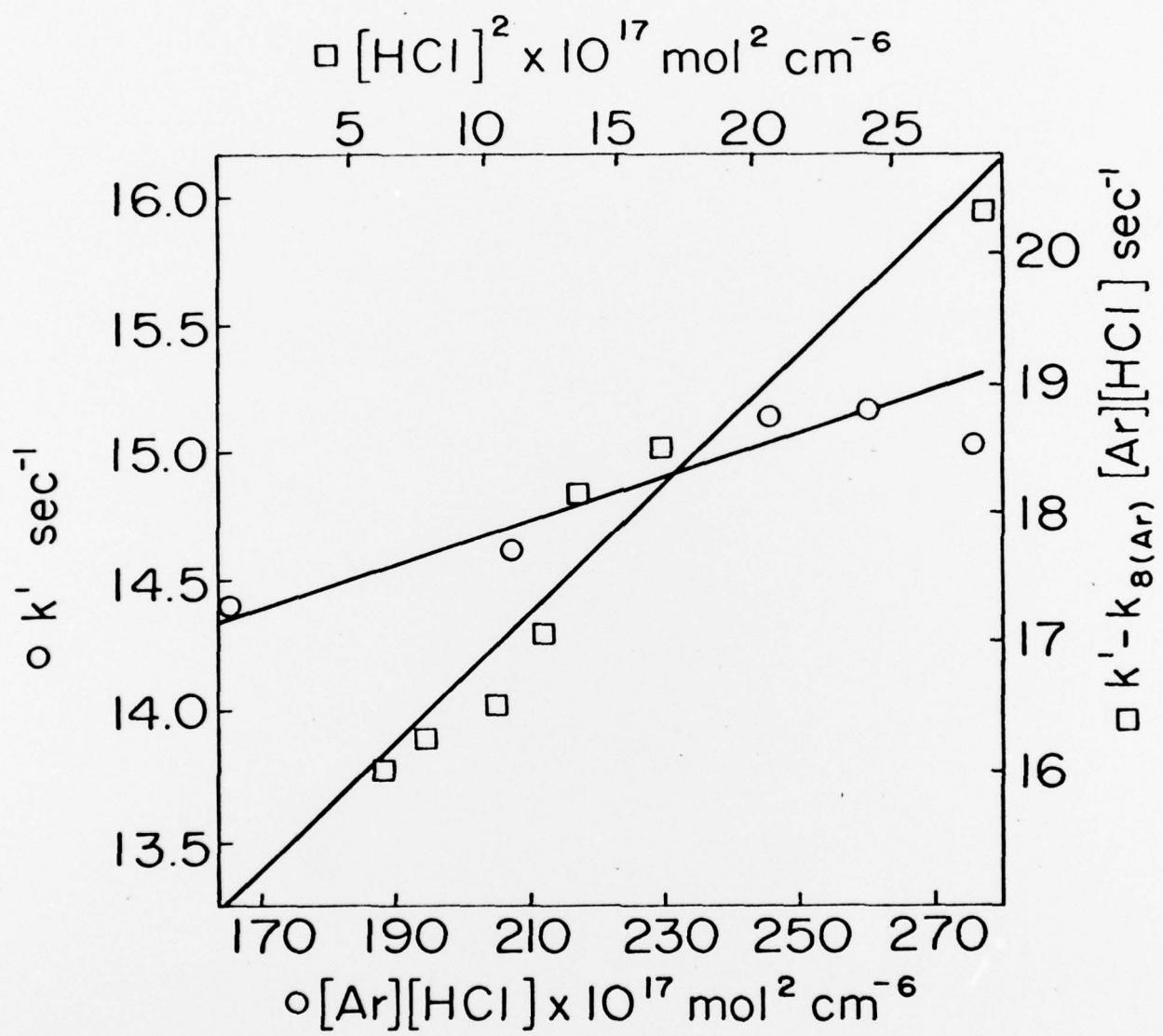


Figure 4

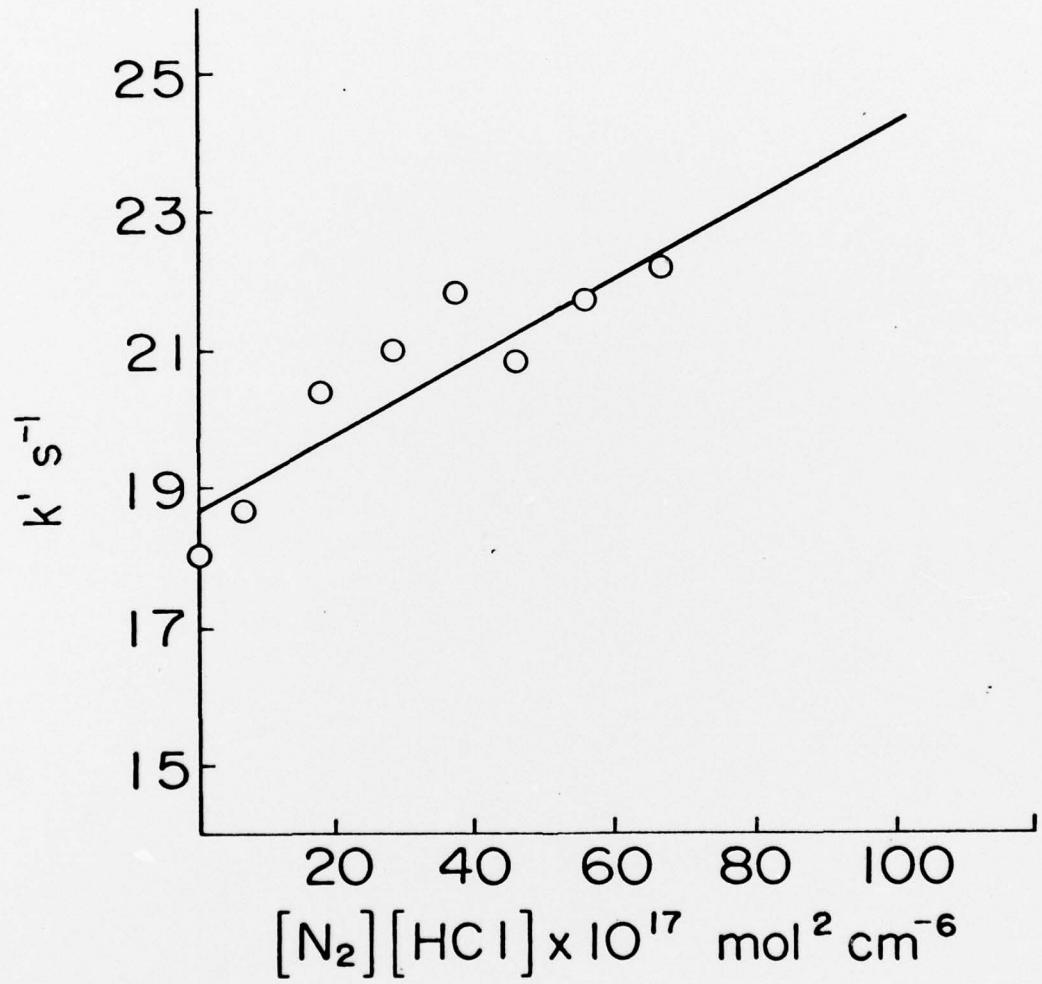


Figure 5